

Supplementary Material**Chiral Coordination Polymers of Mandelate and its Derivatives: Tuning Crystal Packing by Modulation of Hydrogen Bonding***Hui Min Tay^A and Carol Hua^{A,B}*^ASchool of Chemistry, The University of Melbourne, Parkville, Vic. 3010, Australia.^BCorresponding author. Email: carol.hua@unimelb.edu.au**Table of Contents**

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Table S1. Crystallographic parameters compounds **1-3** in this study.

Compound	[Ni(<i>R</i>- man)₂(bpb)(H₂O)] (1)	[Co(<i>R</i>- Man)₂(dpt)(H₂O)] (2)	[Co(<i>S</i>- Man)₂(dptztz)(H₂O)] (3)
Formula	C ₃₂ H ₂₈ N ₂ NiO ₇	C ₂₈ H ₂₄ CoN ₆ O ₇	C ₃₀ H ₂₄ CoN ₄ O ₇ S ₂
Formula Weight	611.27	615.46	675.58
Temperature (K)	100(2)	100(2)	100(2)
Crystal system	monoclinic	monoclinic	triclinic
Space Group	<i>P</i> 2 ₁	<i>P</i> 2 ₁	<i>P</i> 1
<i>a</i> (Å)	5.7090(11)	5.7030(11)	5.6100(11)
<i>b</i> (Å)	15.069(3)	14.430(3)	8.0200(16)
<i>c</i> (Å)	15.578(3)	15.405(3)	16.030(3)
<i>α</i> (°)	90	90	84.22(3)
<i>β</i> (°)	98.51(3)	98.06(3)	80.45(3)
<i>γ</i> (°)	90	90	88.94(3)
Cell Volume (Å³)	1325.4(5)	1255.2(4)	707.6(3)
<i>Z</i>	2	2	1
ρ_{calc}/cm³	1.532	1.628	1.585
μ/mm⁻¹	0.788	0.748	0.811
<i>F</i>(000)	636.0	634.0	347.0
Crystal size/mm³	0.06 × 0.04 × 0.02	0.02 × 0.015 × 0.013	0.12 × 0.08 × 0.08
Radiation	Synchrotron (λ = 0.71073)	Synchrotron (λ = 0.71050)	Synchrotron (λ = 0.71073)
Reflections collected	8637	21810	13593
Independent reflections	4815 [<i>R</i> _{int} = 0.0751, <i>R</i> _{sigma} = 0.0895]	6815 [<i>R</i> _{int} = 0.0388, <i>R</i> _{sigma} = 0.0336]	5688 [<i>R</i> _{int} = 0.0207, <i>R</i> _{sigma} = 0.0235]
Data/restraints /parameters	4815/1/384	6815/4/387	5688/6/396
GooF	1.083	1.037	1.139
<i>R</i>₁, <i>wR</i>₂ (<i>I</i> > 2σ(<i>I</i>))	<i>R</i> ₁ = 0.0839, <i>wR</i> ₂ = 0.2298	<i>R</i> ₁ = 0.0256, <i>wR</i> ₂ = 0.0645	<i>R</i> ₁ = 0.0299, <i>wR</i> ₂ = 0.0751
<i>R</i>₁, <i>wR</i>₂ (all)	<i>R</i> ₁ = 0.0971, <i>wR</i> ₂ = 0.2399	<i>R</i> ₁ = 0.0257, <i>wR</i> ₂ = 0.0646	<i>R</i> ₁ = 0.0327, <i>wR</i> ₂ = 0.0890
Largest diff. peak/hole/eÅ⁻³	2.27/−0.64	0.29/−0.59	0.43/−0.79
Flack parameter	0.03(2)	−0.005(4)	0.005(5)

Table S2. Crystallographic parameters compounds **4-6** in this study.

Compound	[Ni(2,4-difluoro- man)₂(bpe)] (4)	[Ni((<i>R</i>)- ManOMe)₂(bpe)] (5)	[Co((<i>R</i>)- ManOMe)₂(bpe)] (6)
Formula	C ₂₈ H ₂₀ F ₄ N ₂ NiO ₆	C ₃₀ H ₃₀ N ₂ NiO ₆	C ₃₀ H ₃₀ CoN ₂ O ₆
Formula Weight	615.17	573.27	573.49
Temperature (K)	100(2)	100(2)	100(2)
Crystal system	monoclinic	monoclinic	monoclinic
Space Group	<i>C2/c</i>	<i>C2</i>	<i>C2</i>
<i>a</i> (Å)	25.335(5)	14.520(3)	14.570(3)
<i>b</i> (Å)	9.938(2)	15.890(3)	15.830(3)
<i>c</i> (Å)	10.850(2)	16.660(3)	16.760(3)
α (°)	90	90	90
β (°)	101.72(3)	114.03(3)	111.49(3)
γ (°)	90	90	90
Cell Volume (Å³)	2674.8(10)	3510.6(14)	3596.9(14)
<i>Z</i>	4	2	4
ρ_{calc}/cm³	1.528	1.085	1.059
μ/mm⁻¹	0.799	0.589	0.512
F(000)	1256.0	1200.0	1196.0
Crystal size/mm³	0.05 × 0.05 × 0.05	0.15 × 0.12 × 0.02	0.2 × 0.18 × 0.10
Radiation	Synchrotron ($\lambda = 0.710876$)	Synchrotron ($\lambda = 0.71073$)	Synchrotron ($\lambda = 0.71073$)
Reflections collected	23233	47128	48404
Independent reflections	3992 [$R_{\text{int}} = 0.0895$, $R_{\text{sigma}} = 0.0568$]	8279 [$R_{\text{int}} = 0.0387$, $R_{\text{sigma}} = 0.0225$]	8427 [$R_{\text{int}} = 0.0390$, $R_{\text{sigma}} = 0.0242$]
Data/restraints /parameters	3992/129/170	8279/1/305	8427/1/219
GooF	1.038	1.321	1.062
R_1, wR_2 ($I > 2\sigma(I)$)	$R_1 = 0.0773$, $wR_2 = 0.2150$	$R_1 = 0.0846$, $wR_2 = 0.2686$	$R_1 = 0.0876$, $wR_2 = 0.2498$
R_1, wR_2 (all)	$R_1 = 0.0900$, $wR_2 = 0.2327$	$R_1 = 0.0864$, $wR_2 = 0.2743$	$R_1 = 0.0911$, $wR_2 = 0.2543$
Largest diff. peak/hole/eÅ⁻³	1.40/−1.31	1.41/−0.69	1.83/−1.18
Flack parameter	N/A	0.09(2)	0.20(4)

Table S3. Analysis of the possible coordination geometries using the SHAPE program for the 6-coordinate M(II) centres in compounds **1-3**

Geometry	Symmetry	1 (Ni)	2 (Co)	3 (Ni)	4 (Co)
HP-6	D _{6h}	31.691	32.103	31.898	29.950
PPY-6	C _{5v}	27.135	25.191	24.849	28.569
OC-6	O _h	0.464	0.717	0.815	0.455
TPR-6	D _{3h}	14.145	13.292	13.300	15.938
JPPY-6	C _{5v}	30.893	28.867	28.474	31.611

HP-6 = Hexagon; PPY-6 = Pentagonal pyramid; OC-6 = Octahedron; TPR-6 = Trigonal prism; JPPY-6 = Johnson pentagonal pyramid 32. The minima values are indicated in bold.

Table S4. Analysis of the possible coordination geometries using the SHAPE program for the 6-coordinate M(II) centres in compounds **1-3**

Geometry	Symmetry	5 (Ni1)	5 (Ni2)	6 (Co1)	6 (Co2)
HP-6	D _{6h}	31.424	31.989	33.553	32.863
PPY-6	C _{5v}	26.746	26.452	24.614	24.525
OC-6	O _h	0.547	0.751	1.398	1.293
TPR-6	D _{3h}	14.412	12.797	10.934	11.345
JPPY-6	C _{5v}	30.292	30.200	27.862	27.893

HP-6 = Hexagon; PPY-6 = Pentagonal pyramid; OC-6 = Octahedron; TPR-6 = Trigonal prism; JPPY-6 = Johnson pentagonal pyramid 32. The minima values are indicated in bold.

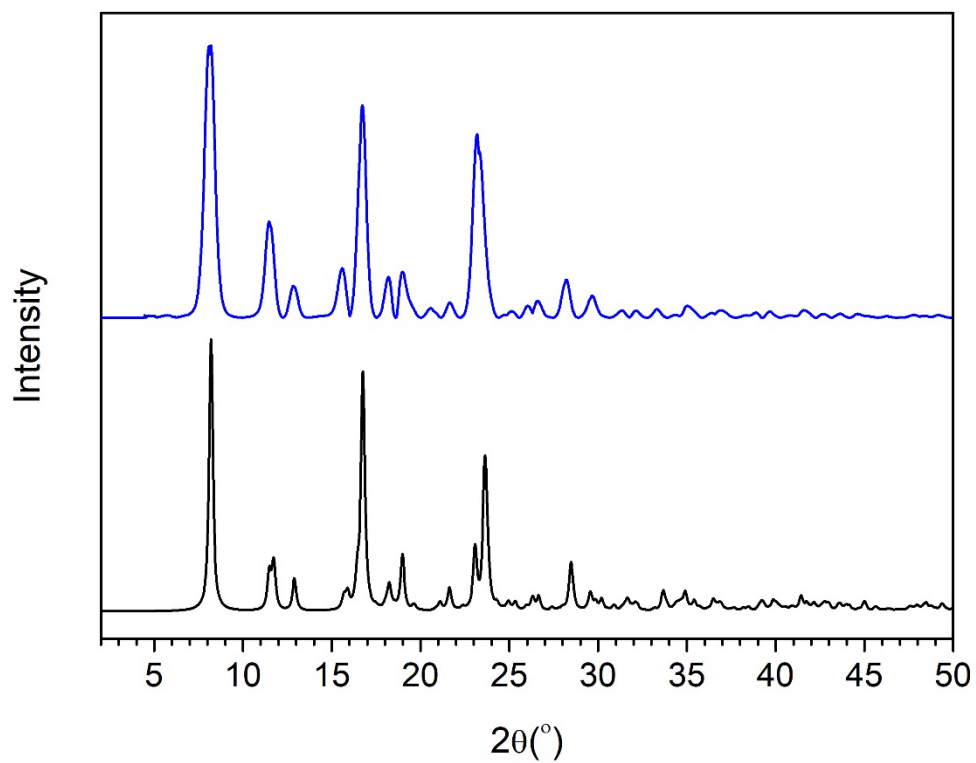


Figure S1. PXRD of [Ni(*R*-Man)₂(bpb)(H₂O)] (1) (blue) between 2 and 50° 2θ versus the calculated pattern (black).

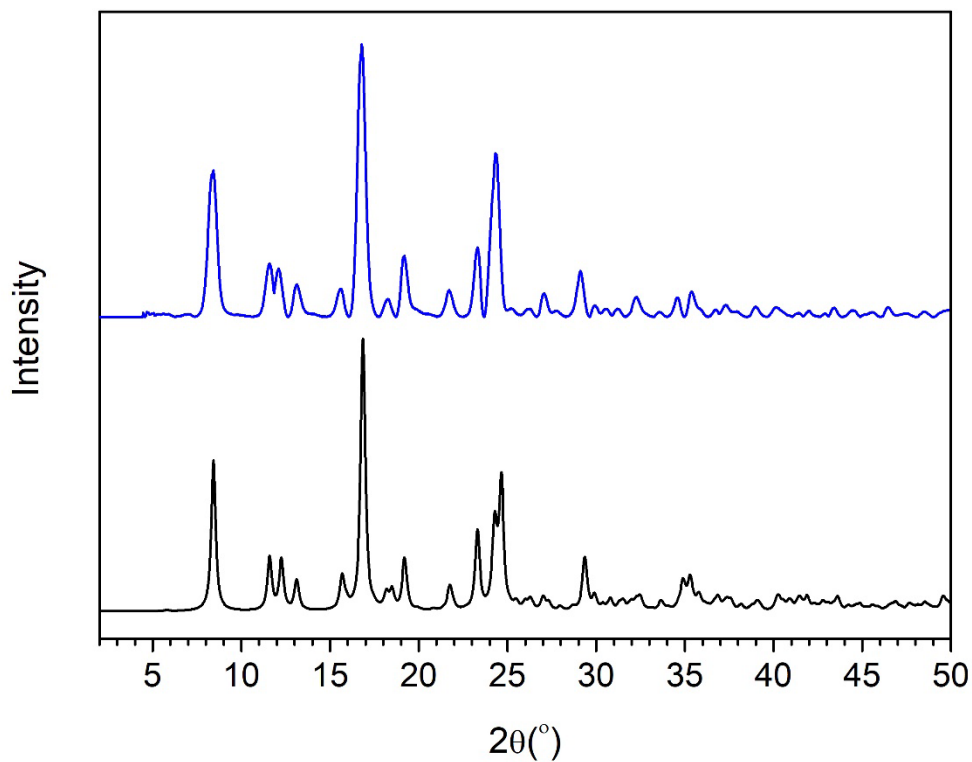


Figure S2. PXRD of [Co(*R*-Man)₂(bpt)(H₂O)] (2) (blue) between 5 and 50° 2θ versus the calculated pattern (black).

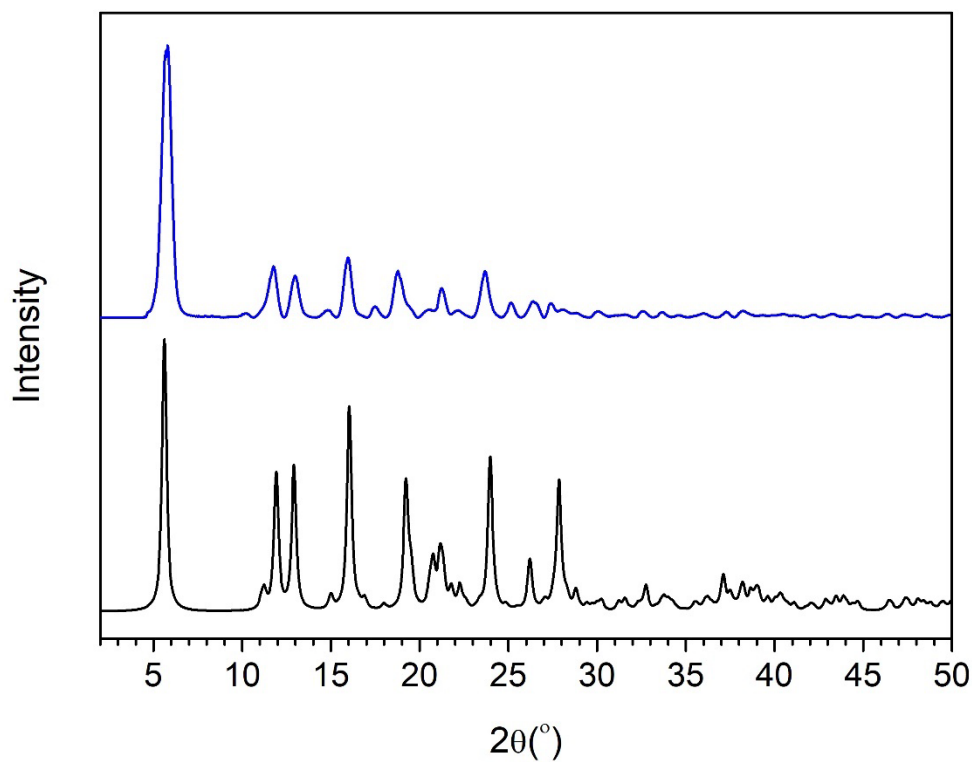


Figure S3. PXRD of $[\text{Co}(\text{S-Man})_2(\text{dptztz})(\text{H}_2\text{O})]$ (**3**) (blue) between 5 and 50° 2θ versus the calculated pattern (black).

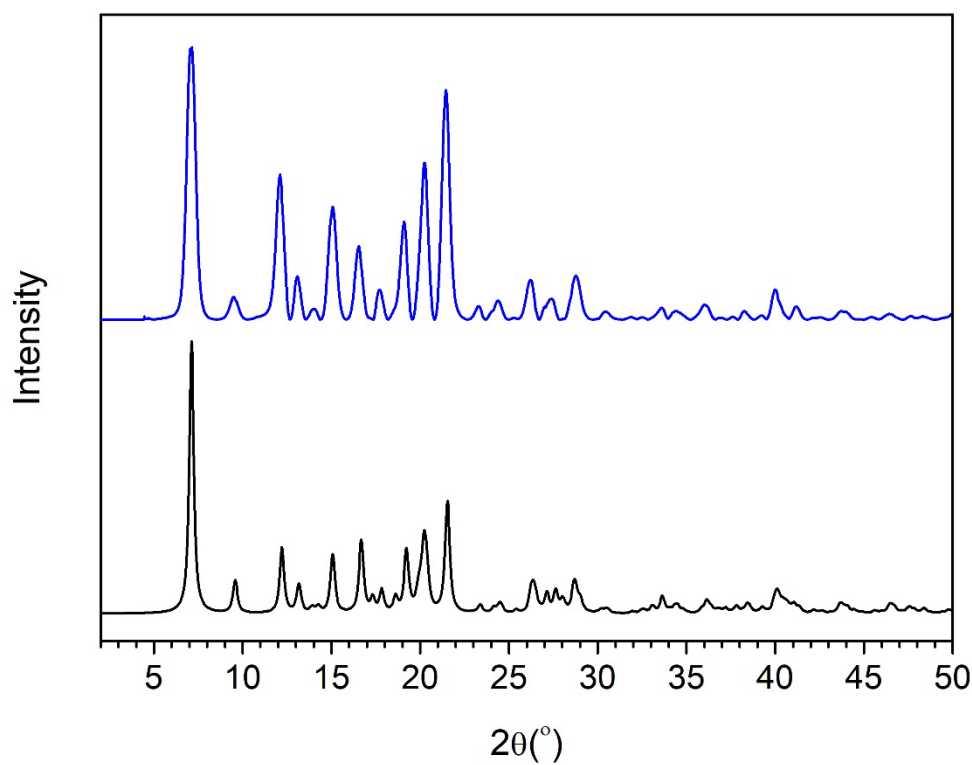


Figure S4. PXRD of $[\text{Ni}(\text{F}_2\text{Man})_2(\text{bpe})]$ (**4**) (blue) between 5 and 50° 2θ versus the calculated pattern (black).

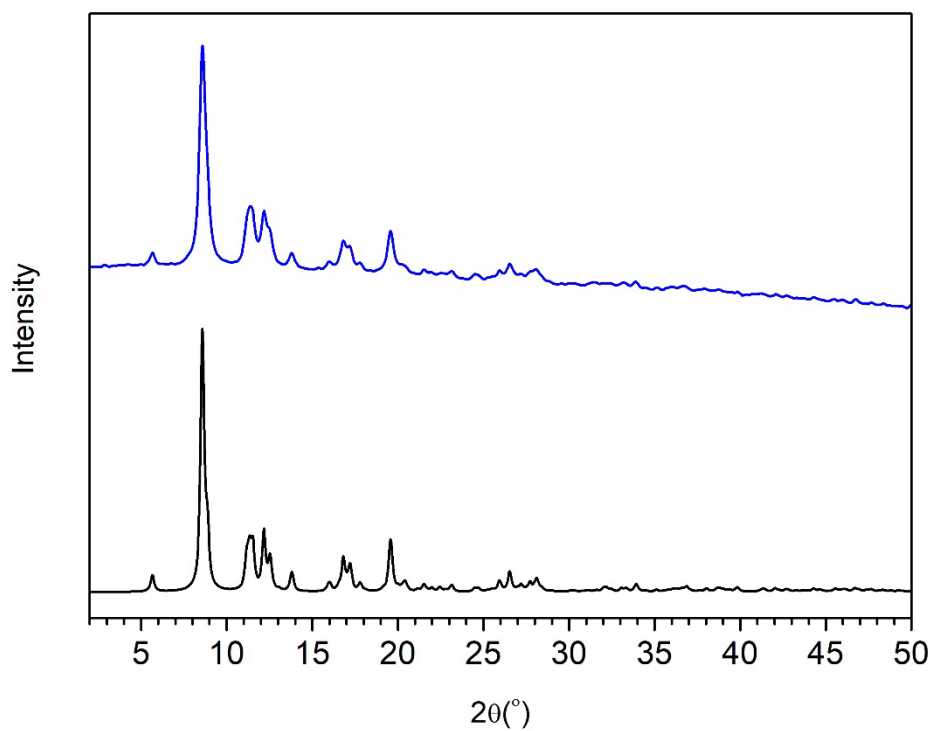


Figure S5. PXRD of $[\text{Ni}(\text{bpe})(R\text{-ManOMe})_2]$ (**5**) (blue) between 5 and 50° 2θ versus the calculated pattern (black).

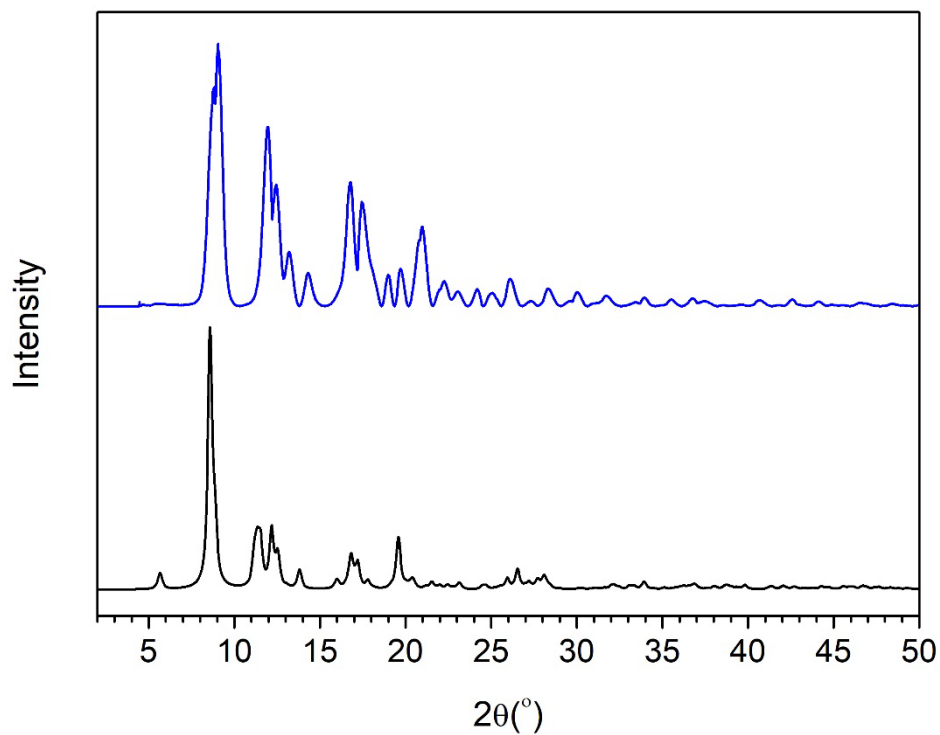


Figure S6. PXRD of $[\text{Co}(\text{bpee})(R\text{-ManOMe})_2]$ (**6**) (blue) between 5 and 50° 2θ versus the calculated pattern (black).

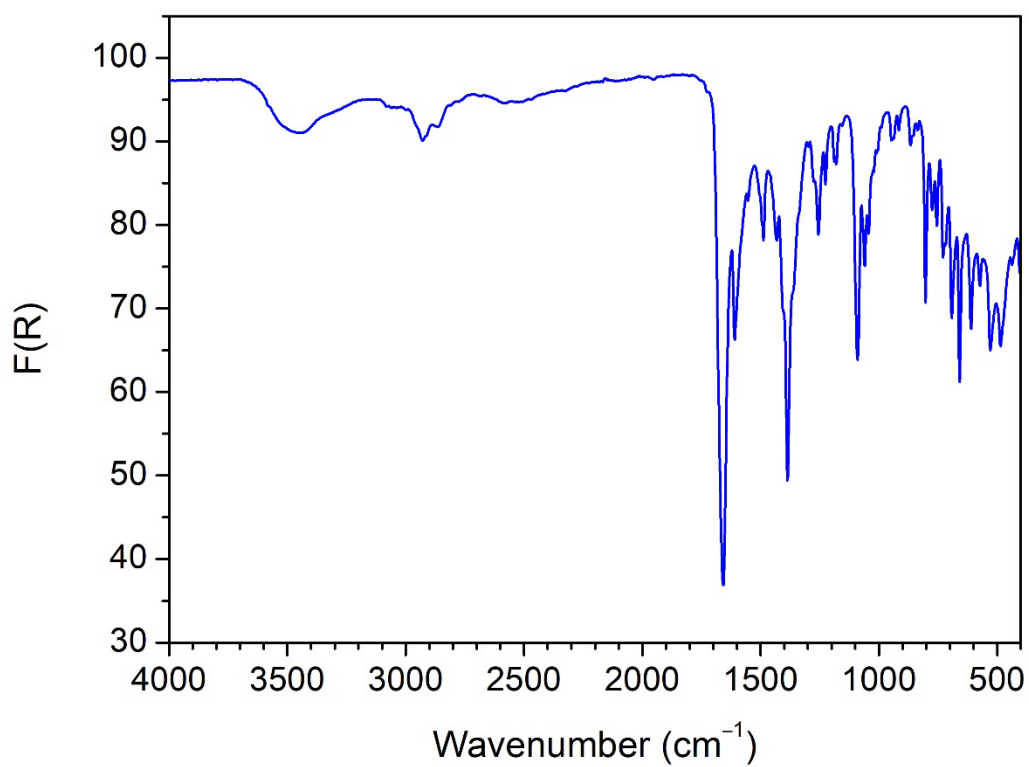


Figure S7. ATR FT-IR spectra of [Ni(*R*-man)₂(bpb)(H₂O)] (**1**) between 4000 and 400 cm⁻¹.

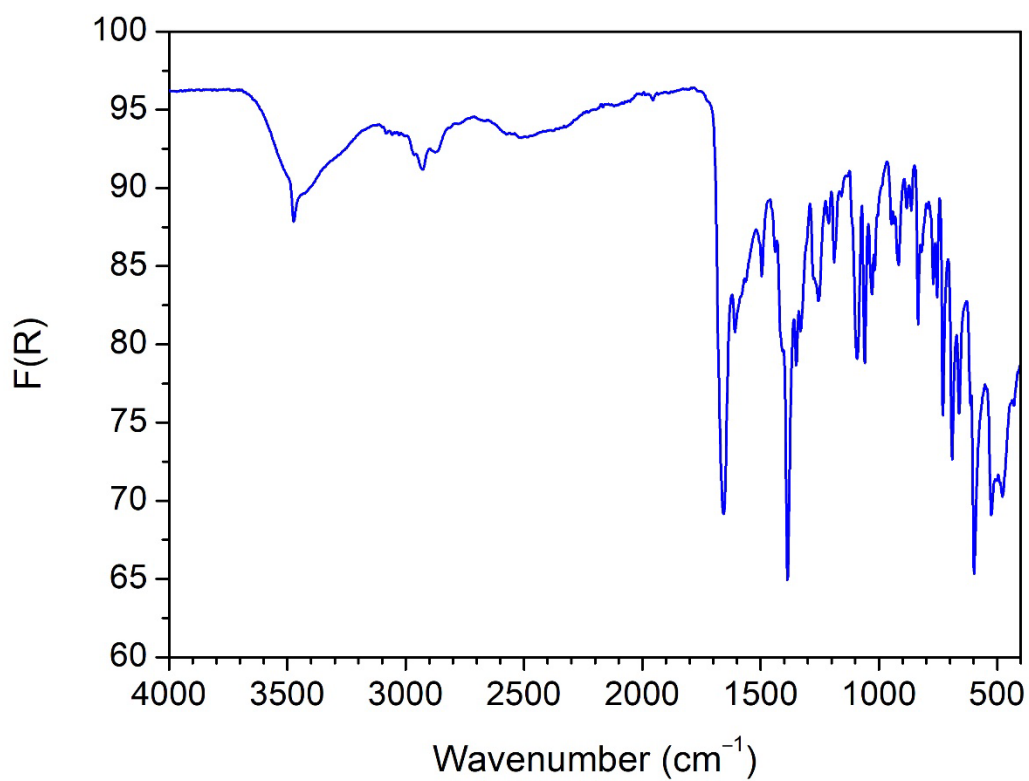


Figure S8. ATR FT-IR spectra of [Co(*R*-Man)₂(bpt)(H₂O)] (**2**) between 4000 and 400 cm⁻¹.

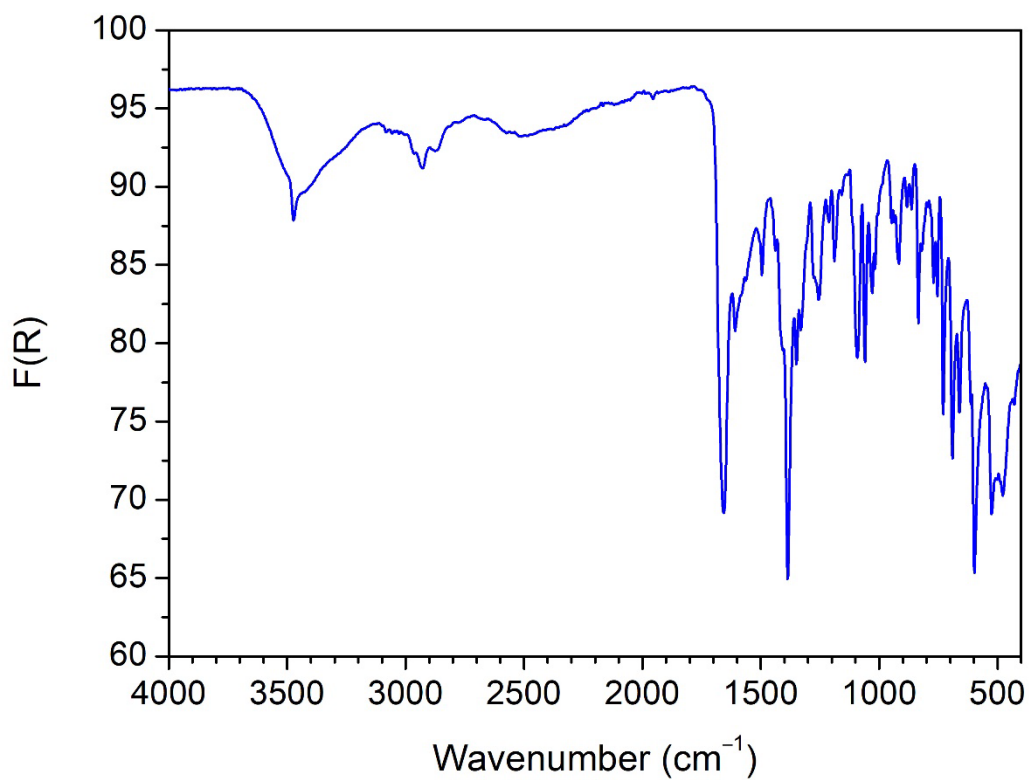


Figure S9. ATR FT-IR spectra of [Co(S-Man)₂(dptztz)(H₂O)] (3) between 4000 and 400 cm⁻¹.

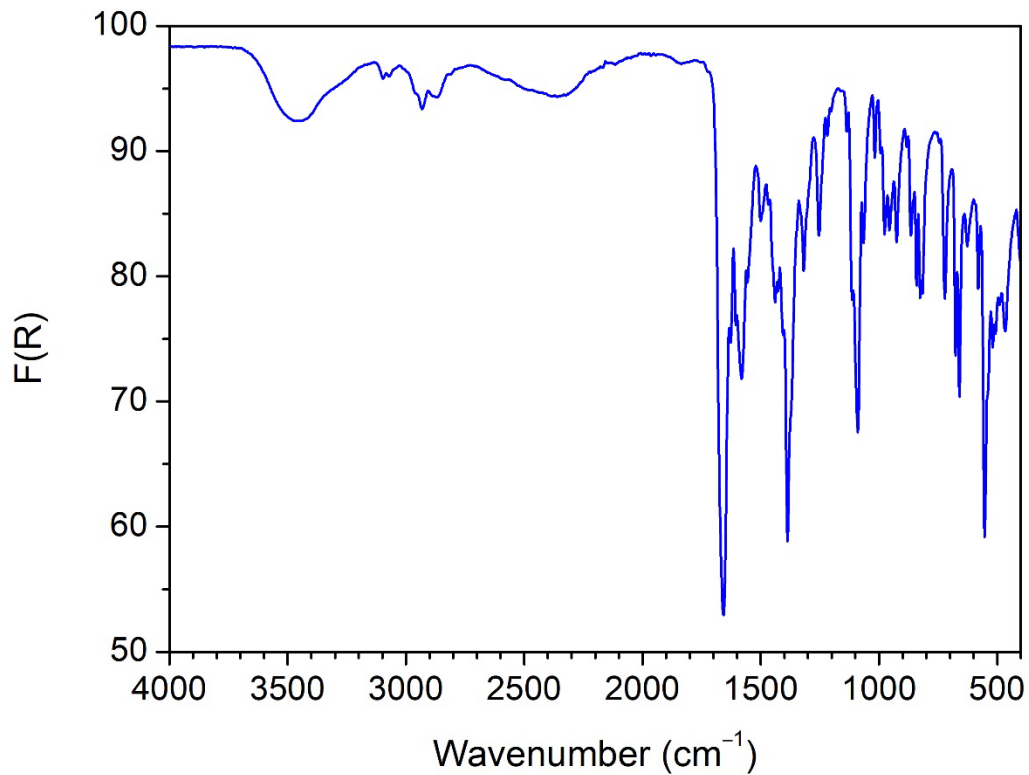


Figure S10. ATR FT-IR spectra of [Ni(F₂Man)₂(bpe)] (4) between 4000 and 400 cm⁻¹.

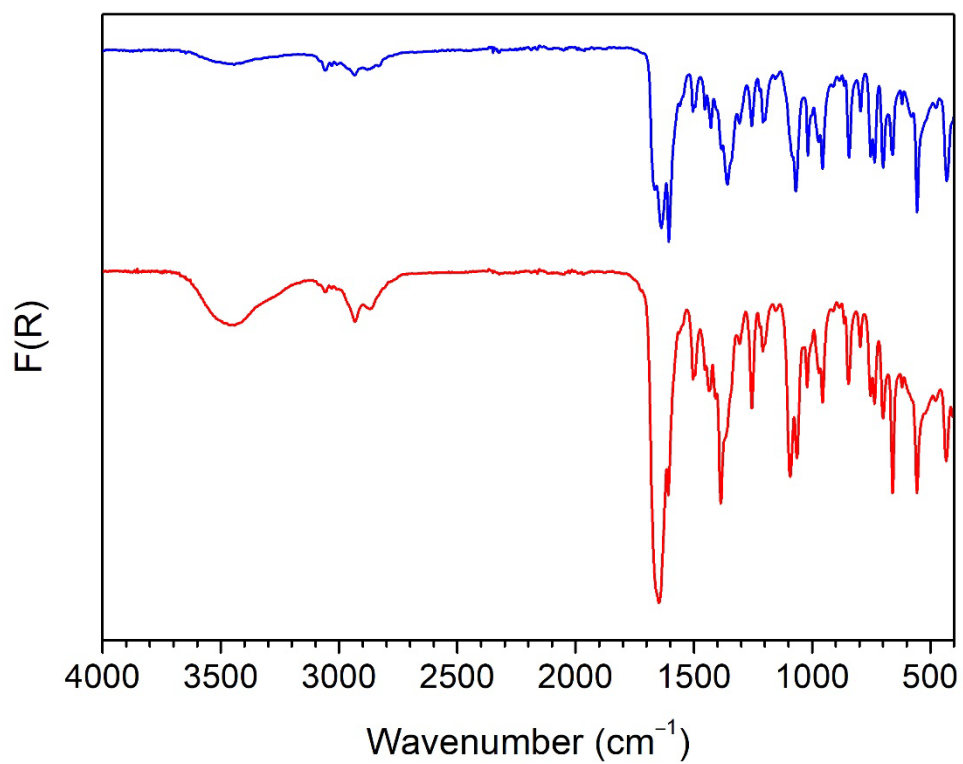


Figure S11. ATR FT-IR spectra of [Ni(bpe)(*R*-ManOMe)₂] (**5**) (red) and [Co(bpee)(*R*-ManOMe)₂] (**6**) (blue) between 4000 and 400 cm⁻¹.

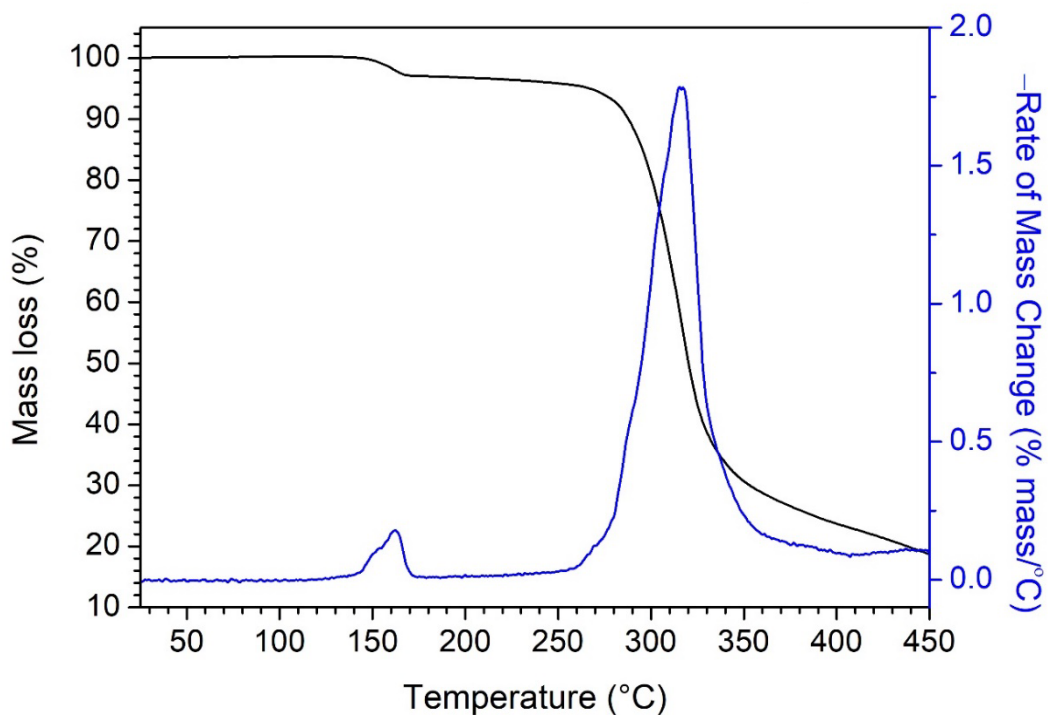


Figure S12. TGA of $[\text{Ni}(\text{R-man})_2(\text{bpb})(\text{H}_2\text{O})]$ (1) between 25 and 400 °C measured under nitrogen gas with a ramp rate of 10 °C/min.

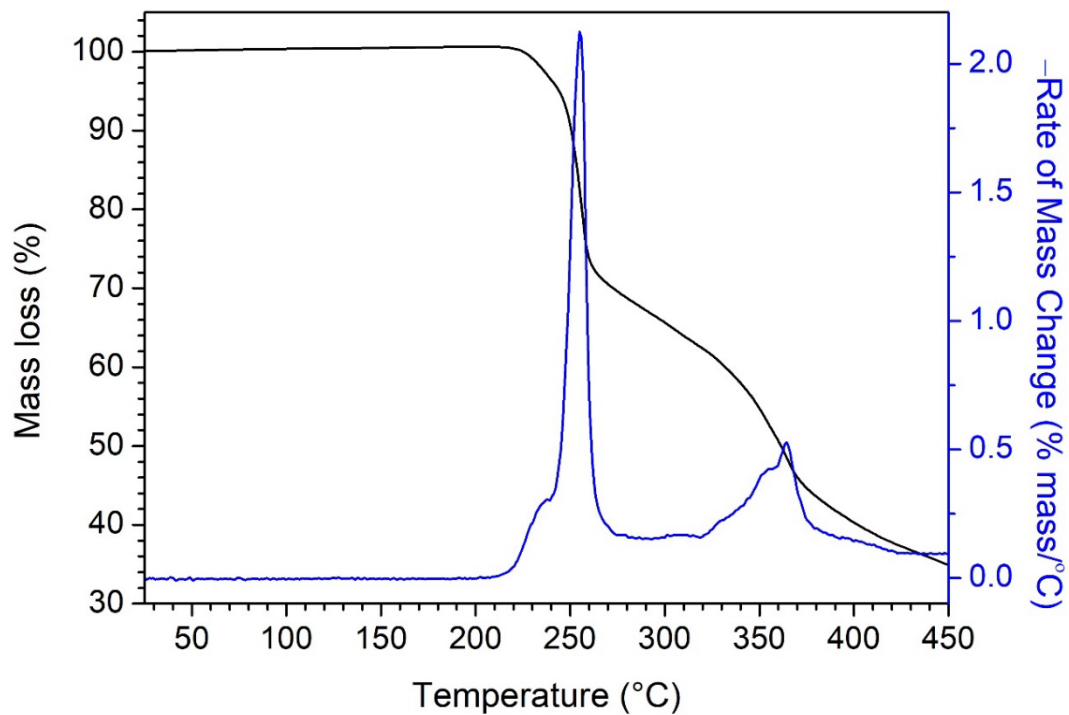


Figure S13. TGA of $[\text{Co}(\text{R-Man})_2(\text{bpt})(\text{H}_2\text{O})]$ (2) between 25 and 400 °C measured under nitrogen gas with a ramp rate of 10 °C/min.

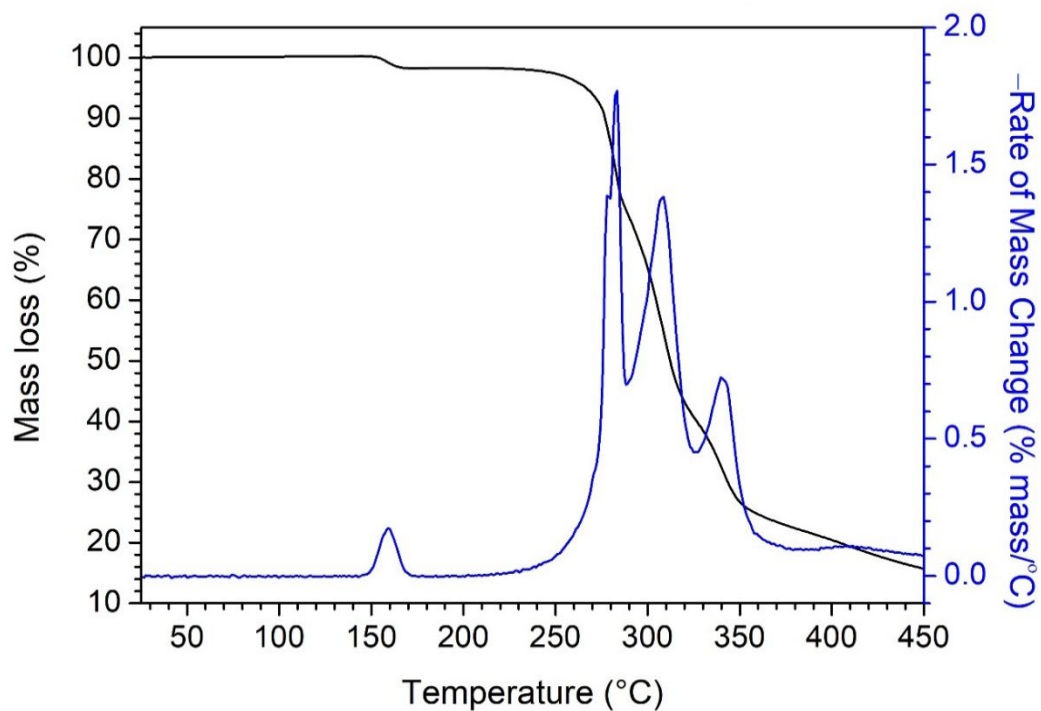


Figure S14. TGA of $[\text{Co}(\text{S-Man})_2(\text{dptztz})(\text{H}_2\text{O})]$ (**3**) between 25 and 400 °C measured under nitrogen gas with a ramp rate of 10 °C/min.

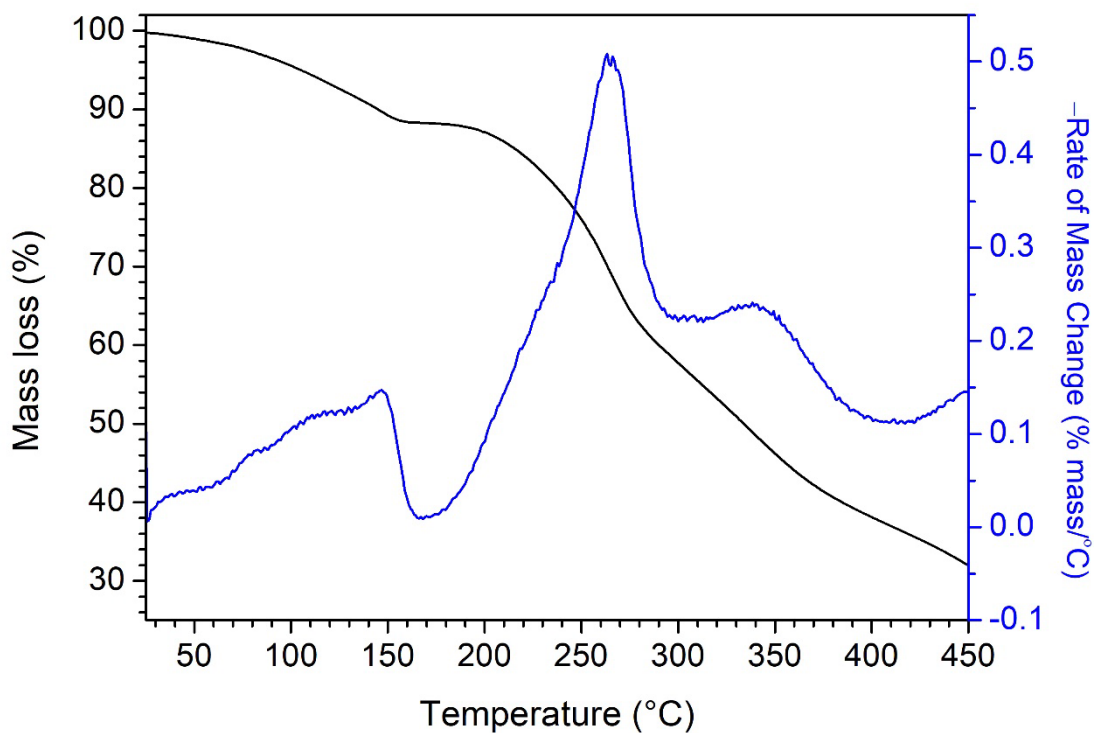


Figure S15. TGA of $[\text{Ni}(\text{F}_2\text{Man})_2(\text{bpe})]$ (**4**) between 25 and 400 °C measured under nitrogen gas with a ramp rate of 10 °C/min.

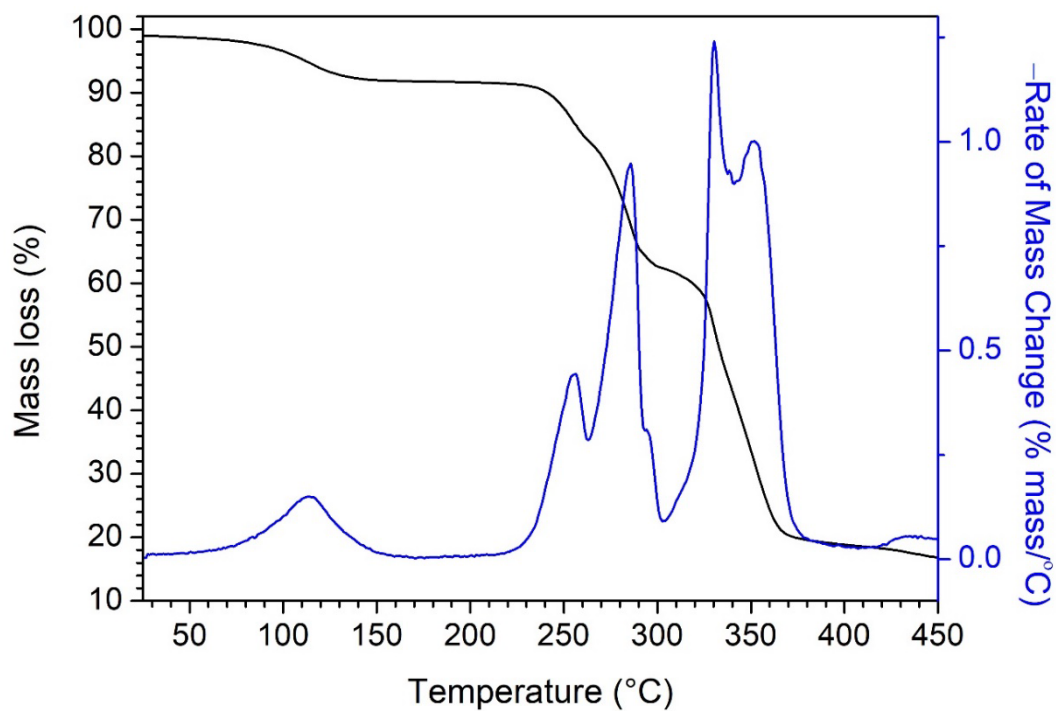


Figure S16. TGA of $[\text{Ni}(\text{bpe})(R\text{-ManOMe})_2]$ (**5**) between 25 and 400 °C measured under nitrogen gas with a ramp rate of 10 °C/min.

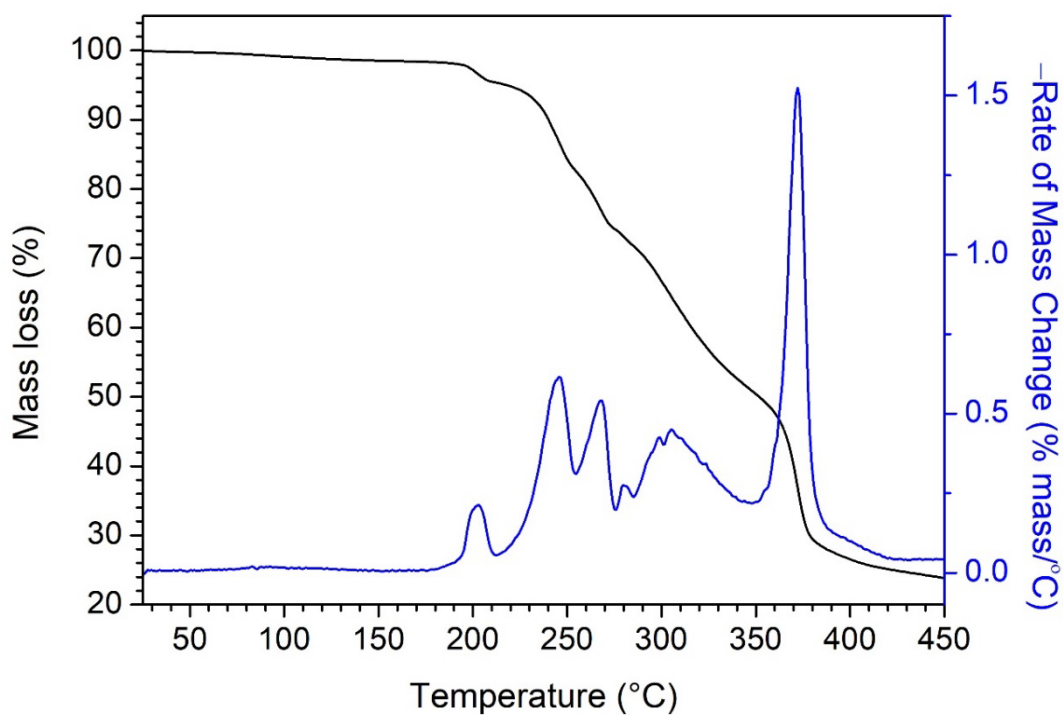


Figure S17. TGA of $[\text{Co}(\text{bpee})(R\text{-ManOMe})_2]$ (**6**) between 25 and 400 °C measured under nitrogen gas with a ramp rate of 10 °C/min.