# THE SOLUBILITY OF 1,4-BUTANESULTONE IN WATER* 

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1,4-Butanesultone (4-hydroxy-1-butanesulphonic acid sultone) has been described as immiscible with water. ${ }^{1}$ However, certain experiments in this laboratory involving reactions of the sultone in aqueous media showed it to be moderately soluble in water. Determination of the mutual solubilities of the sultone (prepared according to ref. ${ }^{2}$, m.p. $13-14^{\circ}$ ) and water over a range of temperatures, using the cloud-point technique and sealed tubes for temperatures above $80^{\circ}$, showed the two liquids to have an upper consolute temperature of $119^{\circ}$ at a composition of $56 \%$ by weight ( 0.144 mole fraction) of 1,4-butanesultone. The consolute temperature is not attainable at atmospheric pressure, since the approximate boiling point at 760 mm of the two-phase mixture was found to be $103^{\circ}$.

Table 1
MUTUAL SOLUBILITY OF WATER AND 1,4 -BUTANESULTONE

| Temp. | Aqueous Phase |  | Sultone Phase |  |
| :---: | :---: | :---: | :---: | :---: |
|  | Weight \% of Sultone | Mole Fraction of Sultone | Weight \% of Sultone | Mole Fraction of Sultone |
| $0^{\circ}$ | $7 \cdot 0$ | $0 \cdot 00986$ |  |  |
| 20 | $7 \cdot 6$ | 0.0108 | $98 \cdot 1$ | 0.872 |
| 40 | $10 \cdot 4$ | $0 \cdot 0151$ | $96 \cdot 4$ | 0.780 |
| 60 | $14 \cdot 6$ | $0 \cdot 0221$ | 93.5 | 0.656 |
| 80 | $23 \cdot 5$ | $0 \cdot 0397$ | $89 \cdot 0$ | 0.517 |
| 90* | 29.7 | $0 \cdot 0529$ | $86 \cdot 0$ | $0 \cdot 448$ |
| 100* | $36 \cdot 5$ | $0 \cdot 0709$ | $79 \cdot 7$ | $0 \cdot 342$ |
| 100* | $45 \cdot 5$ | 0.0955 | $70 \cdot 9$ | $0 \cdot 244$ |
| 119* | 56 |  | $0 \cdot 144$ |  |

* Sealed tube data.

Table 1 shows the mutual solubilities at various temperatures read off from the graph of cloud-point against composition. Hydrolysis of the sultone, measured by acid formation, was not more than $0.1 \%$ in equal volumes of sultone and water held at $130^{\circ}$ for 5 min . Even at the highest temperatures the errors in the quoted solubility data due to hydrolysis are therefore negligible. The degree of hydrolysis found is not incompatible with the data of Helberger and Lantermann ${ }^{1}$ (complete hydrolysis in 90 min at $100^{\circ}$ ) since they studied dilute aqueous solutions.

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[^0]:    * Manuscript received May 19, 1966.
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    ${ }^{1}$ Helberger, J. H., and Lantermann, H., Liebigs Ann., 1954, 586, 147.
    ${ }^{2}$ Snoddy, A. O., Org. Synth., 1963, Coll. Vol. IV, 529.

